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Methyl 2-[(*E*)-(4-nitrophenyl)hydrazono]-3-oxobutyrate

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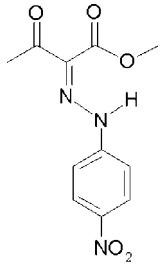
Key indicators: single-crystal X-ray study; T = 296 K; mean σ(C–C) = 0.003 Å; R factor = 0.042; wR factor = 0.121; data-to-parameter ratio = 13.1.

The molecule of the title compound, $C_{11}H_{11}N_3O_5$, exists as the *E* isomer as it is stabilized by an intramolecular hydrogen bond. Except for the methyl H atoms, all atoms lie in special positions on a mirror plane and form a large conjugated system; the methyl H atoms are disordered about the mirror plane. In the crystalline state, bifurcated intra- and intermolecular N—H···O hydrogen bonds and four intermolecular C—H···O hydrogen bonds link the molecules into large perfectly planar sheets. Along the *c* axis, the N—N bond center approaches the phenyl-ring centroids of its neighbouring molecules above and below to give π–π overlap (at a distance of *ca* 3.57 Å), thus fusing the molecules into a three-dimensional framework.

Related literature

For related literature, see: Bernstein *et al.* (1995); Lewis *et al.* (1999); Liu *et al.* (2007, 2008); Mague *et al.* (1997); Mahy *et al.* (1993); Serbutoviez *et al.* (1995); Thami *et al.* (1992); Wang *et al.* (2005).



Experimental

Crystal data

$C_{11}H_{11}N_3O_5$
M_r = 265.2
Orthorhombic, *Pbcm*
a = 12.880 (3) Å
b = 14.299 (3) Å
c = 6.6328 (14) Å
V = 1221.6 (5) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.12 mm^{-1}
T = 296 (2) K
0.30 × 0.30 × 0.20 mm

Data collection

Bruker SMART 1000 CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2002)
T_{min} = 0.966, T_{max} = 0.977
10245 measured reflections
1546 independent reflections
968 reflections with $I > 2\sigma(I)$
R_{int} = 0.041

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.042
$wR(F^2)$ = 0.121
S = 1.03
1546 reflections
118 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.20 e Å^{-3}
$\Delta\rho_{min}$ = −0.13 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N1—H1···O4	0.86	1.98	2.618 (3)	130
C2—H2···O3^{i}	0.93	2.55	3.279 (3)	135
C11—H11*B*···O1^{ii}	0.96	2.57	3.128 (3)	117
N1—H1···O2^{iii}	0.86	2.64	3.439 (3)	154
C5—H5···O2^{iii}	0.93	2.62	3.467 (4)	153
C4^{iii}—H4^{iii}···O4	0.93	2.61	3.518 (3)	167

Symmetry codes: (i) $-x+2, y-\frac{1}{2}, -z+\frac{1}{2}$; (ii) $x, y+1, z$; (iii) $-x+1, y+\frac{1}{2}, -z+\frac{1}{2}$.

Data collection: *SMART* (Bruker, 2002); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *SHELXTL* (Sheldrick, 2008).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CS2083).

References

Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
Bruker, (2002). *SMART*, *SAINT* and *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
Lewis, M., Barnes, C. L., Hathaway, B. A. & Glaser, R. (1999). *Acta Cryst.* C**55**, 975–978.
Liu, X.-L., Zhao, Y., Li, Z.-G. & Liu, Y.-H. (2008). *Acta Cryst.* E**64,** o152.
Liu, Y.-H., Zhao, Y., Liu, X.-L., Tong, B.-W. & Ye, J. (2007). *Acta Cryst.* E**63**, o4072.
Mague, J. T., Vang, S., Berge, D. G. & Wacholtz, W. F. (1997). *Acta Cryst.* C**53**, 973–979.
Mahy, J. P., Gaspard, S. & Mansuy, D. (1993). *Biochemistry*, **32**, 4014–4021.
Serbutoviez, C., Bosshard, C., Knöpfle, G., Wyss, P., Prêtre, P., Gunter, P., Schenk, K., Solari, E. & Chapuis, G. (1995). *Chem. Mater.* **7**, 1198–1206.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.
Thami, T., Bassoul, P., Petit, M., Simon, J., Fort, A., Barzoukas, M. & Villaeys, A. (1992). *J. Am. Chem. Soc.* **114**, 915–921.
Wang, J.-P., Chen, X.-X. & Zhang, Y.-Q. (2005). *Huaxue Yanjiu*, **16**, 29–31.

supplementary materials

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Methyl 2-[(*E*)-(4-nitrophenyl)hydrazono]-3-oxobutyrate

Y.-H. Liu, G.-Y. Sun, J.-F. Liu, J. Ye and X.-L. Liu

Comment

Phenylhydrazone and its derivatives show remarkable stability and high tendency to form non-centrosymmetric crystal packing (Lewis *et al.*, 1999; Mague *et al.*, 1997) and exceptional electronic, bioactive and chemical properties useful for analytic purposes (Mahy *et al.*, 1993), for biological chemistry (Thami *et al.*, 1992) and also for optical materials (Serbutoviez *et al.*, 1995). As a part of our ongoing research (Liu *et al.*, 2007; Liu *et al.*, 2008), the crystal structure of the title compound was solved.

The molecule of the title compound exists in the (*E*)-isomer configuration, not as the generally more stable (*Z*)-isomer (Schemes 1 and 2). The (*E*)-isomer exists here because of the N—H···O intra-molecular hydrogen bond stabilizes it by forming a pseudo-ring *S*(6) (Bernstein *et al.*, 1995) motif (Fig. 1, Table 1 and 2). The N1—C6 bond distance at 1.397 (3) Å is longer than the expected C═N double bond (1.32 Å) but is shorter than a C—N single bond (1.47 Å) because of the classic sp^2-hybrid nitrogen atom, as also found in our earlier work (Liu *et al.*, 2007, 2008). All these effects may help all non-hydrogen atoms to form a perfect plane which coincides with the mirror plane of the space group, less for the hydrogen atoms of the two methyl groups whose six H atoms are disordered over two orientations.

In the crystal packing the molecules are linked into larger perfectly planar sheets *via* by four C—H···O inter-molecular hydrogen bonds and one N—H···O intra-molecular hydrogen bond running parallel to the [001] plane (Fig. 2, Table 2). H1 atom of the N1 atom is a part of a bifurcated system and makes both intra- and intermolecular H-bridges, with angles around the H1 adding up to 360°. Finally, along the *c* axis the N1—N2 bond centers of molecules combine its up and down neighbours' phenyl rings into three dimensional framework (Fig. 2). Consecutive bond centers···phenyl ring centers are at a distance of ca. 3.57 Å and an incline at an angle of ca. 137° (Fig. 3).

Experimental

The title compound was synthesized according to literature procedure (Wang *et al.* 2005; Liu *et al.* 2008). Crystals suitable for X-ray diffraction were obtained by slow evaporation of a solution of the solid in dichloromethane at room temperature over a period of 6 d.

Refinement

After their location in a difference map, all H atoms were fixed geometrically at ideal positions and allowed to ride on the parent C atoms, with C—H distances of 0.93 (aromatic) or 0.97 Å (methyl), and with U_{iso}(H) values of $1.2U_{eq}$ (C, N).

Figures

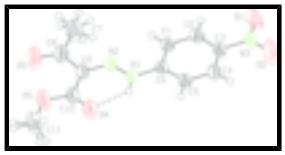

Fig. 1. The molecular structure of the title compound, showing 30% probability displacement ellipsoids. Disorder of the two methyl groups are indicated and the N–H···O intra-molecular hydrogen bond shown as dashed lines.

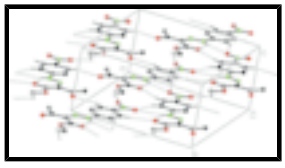

Fig. 2. Part of the crystal structure of the title compound, showing the formation of a hydrogen bonded plane parallel to [001], which is built by one N—H···O and four C—H···O intermolecular hydrogen bonds (dashed lines). For the sake of clarity, H atoms not involved in hydrogen bonding have been omitted.

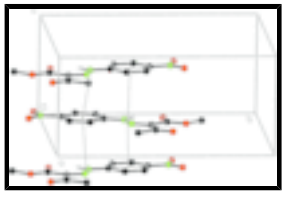

Fig. 3. Excerpt of the crystal structure of the title compound, showing that along the *c* axis the N1—N2 bond center of one molecule combines its up and down phenyl rings in the other two molecules into a three dimensional framework. H atoms not involved in hydrogen bonding have been omitted.

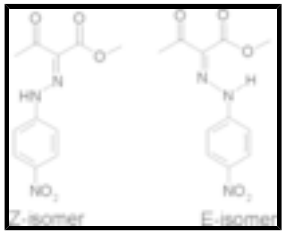

Fig. 4. The *E* and *Z* isomers of the title compound.

(*Z*)-3-Ferrocenyl-2-(4-pyridyl)propenenitrile

Crystal data

$C_{11}H_{11}N_3O_5$	$D_x = 1.442$ Mg m^{-3}
$M_r = 265.2$	Melting point: 400 K
Orthorhombic, *Pbcm*	Mo *K*α radiation λ = 0.71073 Å
Hall symbol: -P 2c 2b	Cell parameters from 1996 reflections
a = 12.880 (3) Å	θ = 2.8–25.4º
b = 14.299 (3) Å	μ = 0.12 mm^{-1}
c = 6.6328 (14) Å	*T* = 296 (2) K
V = 1221.6 (5) Å^3	Block, yellow
Z = 4	0.30 × 0.30 × 0.20 mm
F_{000} = 552	

Data collection

Bruker SMART 1000 CCD diffractometer	1546 independent reflections
Radiation source: fine-focus sealed tube	968 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{int} = 0.041$
T = 296(2) K	$\theta_{max} = 27.6º$

Thin–slice ω scans	θ_{min} = 1.6º
Absorption correction: multi-scan (SADABS; Bruker, 2002)	h = −16→16
T_{min} = 0.966, T_{max} = 0.977	k = −17→18
10245 measured reflections	l = −8→8

Refinement

Refinement on F^2	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)]$ = 0.042	$w = 1/[\sigma^2(F_o^2) + (0.0497P)^2 + 0.2977P]$ where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2)$ = 0.121	$(\Delta/\sigma)_{max} < 0.001$
S = 1.03	$\Delta\rho_{max}$ = 0.20 e Å^{-3}
1546 reflections	$\Delta\rho_{min}$ = −0.13 e Å^{-3}
118 parameters	Extinction correction: SHELXL97 (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Primary atom site location: structure-invariant direct methods	Extinction coefficient: 0.0036 (9)
Secondary atom site location: difference Fourier map	

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}	Occ. (<1)
O4	0.68298 (15)	0.51128 (12)	0.2500	0.0726 (6)	
O3	1.00596 (16)	0.52888 (14)	0.2500	0.0896 (8)	
O1	0.68240 (17)	−0.10594 (13)	0.2500	0.0887 (8)	
O2	0.52310 (17)	−0.06963 (13)	0.2500	0.1081 (10)	
N3	0.61352 (18)	−0.04820 (14)	0.2500	0.0582 (6)	
N1	0.72021 (14)	0.33126 (12)	0.2500	0.0422 (5)	
H1	0.6721	0.3730	0.2500	0.051*	
N2	0.81817 (14)	0.35582 (13)	0.2500	0.0435 (5)	
C10	0.7757 (2)	0.52229 (16)	0.2500	0.0489 (6)	
C6	0.69506 (16)	0.23623 (15)	0.2500	0.0379 (5)	
C8	0.9648 (2)	0.45339 (18)	0.2500	0.0597 (7)	
C5	0.59053 (17)	0.21060 (14)	0.2500	0.0436 (6)	
H5	0.5392	0.2563	0.2500	0.052*	
C2	0.74469 (18)	0.07505 (15)	0.2500	0.0447 (6)	
H2	0.7957	0.0290	0.2500	0.054*	
C4	0.56352 (18)	0.11723 (15)	0.2500	0.0475 (6)	
H4	0.4941	0.0993	0.2500	0.057*	
O5	0.81917 (15)	0.60575 (12)	0.2500	0.0768 (7)	
C3	0.64146 (18)	0.05094 (15)	0.2500	0.0423 (5)	
C1	0.77167 (17)	0.16807 (15)	0.2500	0.0431 (6)	
H1A	0.8413	0.1853	0.2500	0.052*	
C9	0.84922 (18)	0.44293 (16)	0.2500	0.0453 (6)	
C7	1.0288 (2)	0.3663 (2)	0.2500	0.0906 (12)	

H7A	1.0318	0.3412	0.1158	0.136*	0.50
H7B	0.9980	0.3211	0.3387	0.136*	0.50
H7C	1.0978	0.3806	0.2954	0.136*	0.50
C11	0.7477 (3)	0.68338 (19)	0.2500	0.0908 (11)	
H11A	0.7153	0.6882	0.1201	0.136*	0.50
H11B	0.7847	0.7401	0.2788	0.136*	0.50
H11C	0.6955	0.6735	0.3511	0.136*	0.50

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O4	0.0474 (12)	0.0424 (10)	0.1281 (19)	−0.0013 (8)	0.000	0.000
O3	0.0517 (12)	0.0553 (13)	0.162 (2)	−0.0184 (10)	0.000	0.000
O1	0.0783 (14)	0.0355 (10)	0.152 (2)	0.0124 (10)	0.000	0.000
O2	0.0591 (13)	0.0402 (11)	0.225 (3)	−0.0140 (10)	0.000	0.000
N3	0.0587 (14)	0.0335 (11)	0.0824 (16)	−0.0003 (11)	0.000	0.000
N1	0.0398 (10)	0.0318 (10)	0.0551 (12)	−0.0030 (8)	0.000	0.000
N2	0.0430 (11)	0.0392 (10)	0.0482 (12)	−0.0067 (9)	0.000	0.000
C10	0.0524 (16)	0.0355 (13)	0.0588 (16)	−0.0074 (11)	0.000	0.000
C6	0.0426 (12)	0.0333 (11)	0.0377 (12)	−0.0018 (10)	0.000	0.000
C8	0.0493 (15)	0.0478 (15)	0.0821 (19)	−0.0102 (13)	0.000	0.000
C5	0.0417 (12)	0.0317 (12)	0.0575 (14)	0.0039 (9)	0.000	0.000
C2	0.0434 (13)	0.0358 (12)	0.0547 (14)	0.0065 (10)	0.000	0.000
C4	0.0392 (12)	0.0369 (12)	0.0664 (16)	−0.0023 (10)	0.000	0.000
O5	0.0622 (12)	0.0348 (10)	0.1333 (19)	−0.0092 (9)	0.000	0.000
C3	0.0456 (13)	0.0283 (11)	0.0530 (14)	0.0000 (10)	0.000	0.000
C1	0.0381 (12)	0.0404 (13)	0.0507 (14)	−0.0020 (10)	0.000	0.000
C9	0.0460 (13)	0.0360 (12)	0.0537 (14)	−0.0081 (10)	0.000	0.000
C7	0.0511 (17)	0.0568 (17)	0.164 (4)	−0.0011 (14)	0.000	0.000
C11	0.088 (2)	0.0336 (14)	0.150 (3)	−0.0002 (16)	0.000	0.000

Geometric parameters (Å, °)

O4—C10	1.204 (3)	C5—C4	1.380 (3)
O3—C8	1.203 (3)	C5—H5	0.9300
O1—N3	1.212 (3)	C2—C3	1.374 (3)
O2—N3	1.204 (3)	C2—C1	1.375 (3)
N3—C3	1.463 (3)	C2—H2	0.9300
N1—N2	1.310 (2)	C4—C3	1.381 (3)
N1—C6	1.397 (3)	C4—H4	0.9300
N1—H1	0.8600	O5—C11	1.442 (3)
N2—C9	1.308 (3)	C1—H1A	0.9300
C10—O5	1.318 (3)	C7—H7A	0.9600
C10—C9	1.478 (3)	C7—H7B	0.9600
C6—C1	1.387 (3)	C7—H7C	0.9600
C6—C5	1.395 (3)	C11—H11A	0.9600
C8—C7	1.494 (4)	C11—H11B	0.9600
C8—C9	1.496 (3)	C11—H11C	0.9600

O2—N3—O1	122.3 (2)	C3—C4—H4	120.6
O2—N3—C3	119.0 (2)	C10—O5—C11	115.2 (2)
O1—N3—C3	118.7 (2)	C2—C3—C4	122.1 (2)
N2—N1—C6	118.96 (18)	C2—C3—N3	118.8 (2)
N2—N1—H1	120.5	C4—C3—N3	119.1 (2)
C6—N1—H1	120.5	C2—C1—C6	120.0 (2)
C9—N2—N1	123.4 (2)	C2—C1—H1A	120.0
O4—C10—O5	122.7 (2)	C6—C1—H1A	120.0
O4—C10—C9	122.3 (2)	N2—C9—C10	122.4 (2)
O5—C10—C9	115.0 (2)	N2—C9—C8	113.5 (2)
C1—C6—C5	120.1 (2)	C10—C9—C8	124.1 (2)
C1—C6—N1	121.24 (19)	C8—C7—H7A	109.5
C5—C6—N1	118.64 (19)	C8—C7—H7B	109.5
O3—C8—C7	120.3 (2)	H7A—C7—H7B	109.5
O3—C8—C9	121.9 (2)	C8—C7—H7C	109.5
C7—C8—C9	117.8 (2)	H7A—C7—H7C	109.5
C4—C5—C6	119.8 (2)	H7B—C7—H7C	109.5
C4—C5—H5	120.1	O5—C11—H11A	109.5
C6—C5—H5	120.1	O5—C11—H11B	109.5
C3—C2—C1	119.2 (2)	H11A—C11—H11B	109.5
C3—C2—H2	120.4	O5—C11—H11C	109.5
C1—C2—H2	120.4	H11A—C11—H11C	109.5
C5—C4—C3	118.8 (2)	H11B—C11—H11C	109.5
C5—C4—H4	120.6		

C6—N1—N2—C9	180.0	O1—N3—C3—C4	180.0
N2—N1—C6—C1	0.0	C3—C2—C1—C6	0.0
N2—N1—C6—C5	180.0	C5—C6—C1—C2	0.0
C1—C6—C5—C4	0.0	N1—C6—C1—C2	180.0
N1—C6—C5—C4	180.0	N1—N2—C9—C10	0.0
C6—C5—C4—C3	0.0	N1—N2—C9—C8	180.0
O4—C10—O5—C11	0.0	O4—C10—C9—N2	0.0
C9—C10—O5—C11	180.0	O5—C10—C9—N2	180.0
C1—C2—C3—C4	0.0	O4—C10—C9—C8	180.0
C1—C2—C3—N3	180.0	O5—C10—C9—C8	0.0
C5—C4—C3—C2	0.0	O3—C8—C9—N2	180.0
C5—C4—C3—N3	180.0	C7—C8—C9—N2	0.0
O2—N3—C3—C2	180.0	O3—C8—C9—C10	0.0
O1—N3—C3—C2	0.0	C7—C8—C9—C10	180.0
O2—N3—C3—C4	0.0		

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1···O4	0.86	1.98	2.618 (3)	130
C2—H2···O3^{i}	0.93	2.55	3.279 (3)	135
C11—H11B···O1^{ii}	0.96	2.57	3.128 (3)	117
N1—H1···O2^{iii}	0.86	2.64	3.439 (3)	154
C5—H5···O2^{iii}	0.93	2.62	3.467 (4)	153

C4^{iii}—H4^{iii}···O4	0.93	2.61	3.518 (3)	167

Symmetry codes: (i) −*x*+2, *y*−1/2, −*z*+1/2; (ii) *x*, *y*+1, *z*; (iii) −*x*+1, *y*+1/2, −*z*+1/2.

Fig. 1

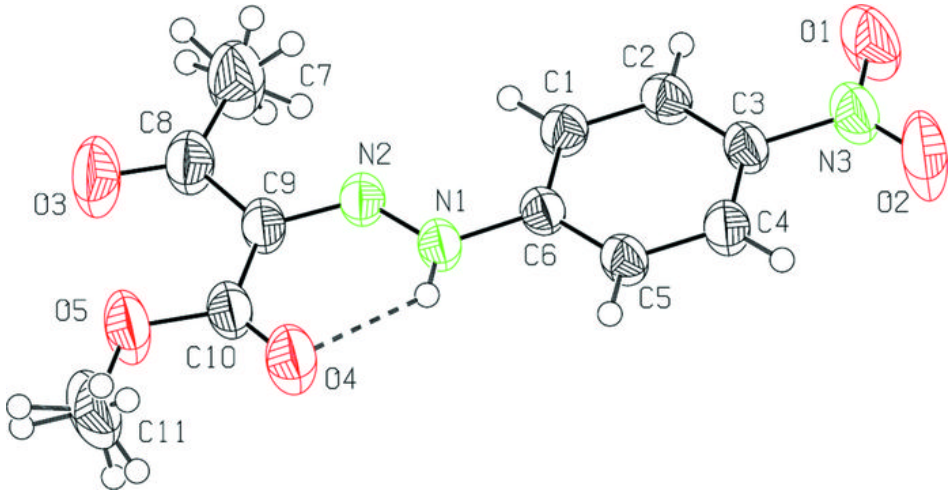

Fig. 2

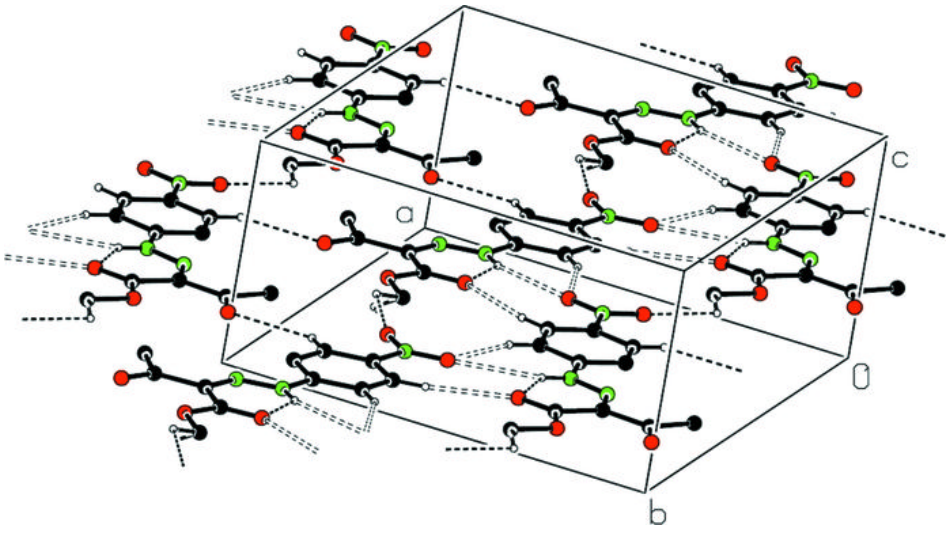

Fig. 3

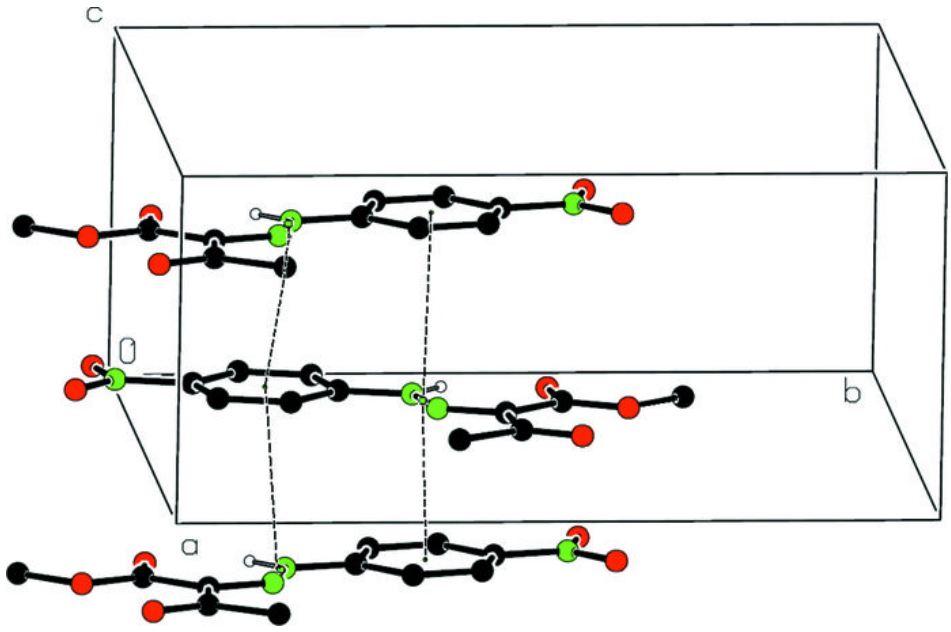

Fig. 4